

MOLAR MASSES PRACTICE #1

Name each substance and determine the molar masses of the following:

- | | FORMULA | MOLAR MASS |
|-----|--|---|
| 1) | Carbon Monoxide <u>CO</u> | $12 + 16 = 28 \text{ g/mol}$ |
| 2) | Chlorine <u>Cl₂</u> | $2(35.4) = 70.8 \text{ g/mol}$ |
| 3) | Potassium Hydroxide <u>KOH</u> | $39 + 16 + 1 = 56 \text{ g/mol}$ |
| 4) | Beryllium Chloride <u>BeCl₂</u> | $9 + 70.8 = 79.8 \text{ g/mol}$ |
| 5) | iron (III) Chloride <u>FeCl₃</u> | $55.8 + 35.4(3) = 162.6 \text{ g/mol}$ |
| 6) | Boron Trifluoride <u>BF₃</u> | $10.8 + 19(3) = 67.8 \text{ g/mol}$ |
| 7) | Magnesium Hydroxide <u>Mg(OH)₂</u> | $24.3 + 2(17.2) = 58.7 \text{ g/mol}$ |
| 8) | Sulfur Dioxide <u>SO₂</u> | $32 + 32 = 64 \text{ g/mol}$ |
| 9) | Ammonium Sulfate <u>(NH₄)₂(SO₄)₂</u> | $18(2) + 96 = 132 \text{ g/mol}$ |
| 10) | Lead (II) Nitrate <u>Pb(NO₃)₂</u> | $207 + 2(62) = 331 \text{ g/mol}$ |
| 11) | Gallium Sulfite <u>Ga₂(SO₃)₃</u> | $69.7(2) + 3(80) = 379.4 \text{ g/mol}$ |
| 12) | lead (II) oxide <u>PbO</u> | $207 + 16 = 223 \text{ g/mol}$ |
| 13) | tin (IV) oxide <u>SnO₂</u> | $118.7 + 32 = 150.7 \text{ g/mol}$ |